**Kinetics and Collision Theory Practice Worksheet**

**FIRST! *IF you need some help remembering activation energy,* watch this video:** [**https://www.youtube.com/watch?v=YacsIU97OFc**](https://www.youtube.com/watch?v=YacsIU97OFc)

**Then, answer these questions:**

1. How does a catalyst speed up a reaction?
2. Use collision theory to explain why reactions should occur more slowly as lower temperatures.
3. Explain how increasing the concentration of reactants would speed up a reaction
4. What does the activation energy of a reaction represent?
5. How is the activation energy related to whether or not a collision between molecules is successful?
6. Why does increasing the temperature speed up the rate of a chemical reaction?
7. What must take place in order for a chemical reaction to happen?
8. Draw a reaction coordinate and label the activation energy, reactants, and products

**Now that you’ve finished answering each of these questions, go back to my weebly and find the answer key to this worksheet. CHECK YOUR WORK!! Make sure you have a solid understanding of these concepts.**