Intermolecular Forces

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| ***Teacher:*** | Ms. Athwal  | ***Date:*** | Nov 12 - Nov 15 | ***Course:*** | Chemistry | ***Grade:*** | 11 |
| ***CA Standard(s):*** 2d *Students know* the atoms and molecules in liquids move in a random pattern relative to one another because the intermolecular forces are too weak to hold the atoms or molecules in a solid form.6b *Students know* how to describe the dissolving process at the molecular level by using the concept of random molecular motion. |
| ***Learning Objective (s):*** LT 4.1 – I can discuss how ionic solids dissolve due to the random molecular motion of the water molecules.LT 4.2 – I can identify a compound as containing hydrogen bonding, dipole – dipole forces, and Van der Waals forces.LT 4.3 – I can compare and contrast various intermolecular forces and relate them to the dissolving process of a solid. |
| ***Essential Question(s):*** How did Fritz Haber both help and hurt billions of people? |
| **Assessment**: * Homework 4.1
* Unit 3 Exam Reflection
* Three question exit ticker on intermolecular forces
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| * ***Do Now***:

1. Fill out the Unit 3 exam reflection. Turn into bin when completed. |

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| **WHOLE GROUP/ DIRECT INTRUCTION** |
| * Intermolecular forces: Dipole-Diple, Van Der Walls Forces, Hydrogen Bonding
* Energy states of molecules in solids, liquids and gasses
* Dissolving factors
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| **SMALL GROUP STATION** |  | **COLLABORATIVE STATION** |  | **COMPUTER ASSISTED STATION** |
| Grading and Reflectingn on Unit 3 ExamFritz Haber literacy reading and annotation  |  | Penny and water drops lab Unit four vocabulary search |  | Watch intermolecular foeces video and take scaffolded notes on powerpoint  |