Intermolecular Forces

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| ***Teacher:*** | Ms. Athwal | ***Date:*** | Nov 12 - Nov 15 | ***Course:*** | Chemistry | ***Grade:*** | 11 |
| ***CA Standard(s):***  2d *Students know* the atoms and molecules in liquids move in a random pattern relative to one another because the intermolecular forces are too weak to hold the atoms or molecules in a solid form.  6b *Students know* how to describe the dissolving process at the molecular level by using the concept of random molecular motion. | | | | | | | |
| ***Learning Objective (s):***  LT 4.1 – I can discuss how ionic solids dissolve due to the random molecular motion of the water molecules.  LT 4.2 – I can identify a compound as containing hydrogen bonding, dipole – dipole forces, and Van der Waals forces.  LT 4.3 – I can compare and contrast various intermolecular forces and relate them to the dissolving process of a solid. | | | | | | | |
| ***Essential Question(s):*** How did Fritz Haber both help and hurt billions of people? | | | | | | | |
| **Assessment**:   * Homework 4.1 * Unit 3 Exam Reflection * Three question exit ticker on intermolecular forces | | | | | | | |
| * ***Do Now***:   1. Fill out the Unit 3 exam reflection. Turn into bin when completed. | | | | | | | |

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| **WHOLE GROUP/ DIRECT INTRUCTION** |
| * Intermolecular forces: Dipole-Diple, Van Der Walls Forces, Hydrogen Bonding * Energy states of molecules in solids, liquids and gasses * Dissolving factors |

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| **SMALL GROUP STATION** |  | **COLLABORATIVE STATION** |  | **COMPUTER ASSISTED STATION** |
| Grading and Reflectingn on Unit 3 Exam  Fritz Haber literacy reading and annotation |  | Penny and water drops lab  Unit four vocabulary search |  | Watch intermolecular foeces video and take scaffolded notes on powerpoint |