Percent Yield and Unit 4 Exam

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| ***Teacher:*** | Ms. Athwal  | ***Date:*** | Dec 9- Dec 13 | ***Course:*** | Chemistry | ***Grade:*** | 11 |
| ***CA Standard(s):*** 3e *Students know* how to calculate the masses of reactants and products in a chemical reaction from the mass of one of the reactants or products and the relevant atomic masses.3f Students know how to calculate percent yield in a chemical reaction. |
| ***Learning Objective (s):*** LT 4.6 – I can convert from grams of one compound to grams of another compound.LT 4.7 – I can convert from particles of one compound to particles of another compoundLT 4.8 – I can complete a complex stoichiometric conversion that incorporates molarity, particles, mass, moles, and volumes of substances.LT 4.9 – I can hypothesize and calculate the percent yield for a given chemical reaction. |
| ***Essential Question(s):*** How much of each reactant do I need for a given amount on product?  |
| **Assessment**: * Unit 4 Exam
* Percent Yield Exit Ticket
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| * ***Do Now***:
1. You run the Haber – Bosh Reaction to make NH3:

N2 + 3H2 🡪 2NH3If you begin with 15 x 1023 particles of H2, then how many grams of NH3 do you produce? |

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| **WHOLE GROUP/ DIRECT INTRUCTION** |
| * Percent Yield
* Stations Review
* Unit 4 Exam
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| **SMALL GROUP STATION** |  | **COLLABORATIVE STATION** |  | **COMPUTER ASSISTED STATION** |
| Unit 4 Remediation  |  | Stations Review  |  | Stations Review  |