Percent Yield and Unit 4 Exam

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| ***Teacher:*** | Ms. Athwal | ***Date:*** | Dec 9- Dec 13 | ***Course:*** | Chemistry | ***Grade:*** | 11 |
| ***CA Standard(s):***  3e *Students know* how to calculate the masses of reactants and products in a chemical reaction from the mass of one of the reactants or products and the relevant atomic masses.  3f Students know how to calculate percent yield in a chemical reaction. | | | | | | | |
| ***Learning Objective (s):***  LT 4.6 – I can convert from grams of one compound to grams of another compound.  LT 4.7 – I can convert from particles of one compound to particles of another compound  LT 4.8 – I can complete a complex stoichiometric conversion that incorporates molarity, particles, mass, moles, and volumes of substances.  LT 4.9 – I can hypothesize and calculate the percent yield for a given chemical reaction. | | | | | | | |
| ***Essential Question(s):*** How much of each reactant do I need for a given amount on product? | | | | | | | |
| **Assessment**:   * Unit 4 Exam * Percent Yield Exit Ticket | | | | | | | |
| * ***Do Now***:  1. You run the Haber – Bosh Reaction to make NH3:   N2 + 3H2 🡪 2NH3  If you begin with 15 x 1023 particles of H2, then how many grams of NH3 do you produce? | | | | | | | |

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| **WHOLE GROUP/ DIRECT INTRUCTION** |
| * Percent Yield * Stations Review * Unit 4 Exam |

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| **SMALL GROUP STATION** |  | **COLLABORATIVE STATION** |  | **COMPUTER ASSISTED STATION** |
| Unit 4 Remediation |  | Stations Review |  | Stations Review |