Complex Stoichiometry

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| ***Teacher:*** | Ms. Athwal | ***Date:*** | Dec 2- Dec 6 | ***Course:*** | Chemistry | ***Grade:*** | 11 |
| ***CA Standard(s):***  6c *Students know* temperature, pressure, and surface area affect the dissolving process.  3e *Students know* how to calculate the masses of reactants and products in a chemical reaction from the mass of one of the reactants or products and the relevant atomic masses. | | | | | | | |
| ***Learning Objective (s):***  LT 4.6 – I can convert from grams of one compound to grams of another compound.  LT 4.7 – I can convert from particles of one compound to particles of another compound  LT 4.8 – I can complete a complex stoichiometric conversion that incorporates molarity, particles, mass, moles, and volumes of substances.  LT 4.9 – I can hypothesize and calculate the percent yield for a given chemical reaction. | | | | | | | |
| ***Essential Question(s):*** Why does soda fizz? | | | | | | | |
| **Assessment**:   * Complex stoichiometry exit ticket * Collaborative Poster Rubric * Percent Yield Exit Ticket | | | | | | | |
| * ***Do Now***:  1. Draw and label what salt dissolving in water looks like 2. What is the molar mass of H to H2O 3. Given the following balanced equation: 8H + 2O2 🡪 4H2O 4. How many grams of Cl are produced from Salt given the following balanced equation: 4NaCl 🡪 4Na +2Cl2 | | | | | | | |

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| **WHOLE GROUP/ DIRECT INTRUCTION** |
| * Complex Stoichiometry grams to grams * Mole Ratios * Percent Yield |

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| **SMALL GROUP STATION** |  | **COLLABORATIVE STATION** |  | **COMPUTER ASSISTED STATION** |
| Practice Unit 3 moles to grams and grams to moles |  | Around the world stations questions |  | Notes on grams to grams complex stoichiometry |