Complex Stoichiometry

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| ***Teacher:*** | Ms. Athwal  | ***Date:*** | Dec 2- Dec 6 | ***Course:*** | Chemistry | ***Grade:*** | 11 |
| ***CA Standard(s):*** 6c *Students know* temperature, pressure, and surface area affect the dissolving process.3e *Students know* how to calculate the masses of reactants and products in a chemical reaction from the mass of one of the reactants or products and the relevant atomic masses. |
| ***Learning Objective (s):*** LT 4.6 – I can convert from grams of one compound to grams of another compound.LT 4.7 – I can convert from particles of one compound to particles of another compoundLT 4.8 – I can complete a complex stoichiometric conversion that incorporates molarity, particles, mass, moles, and volumes of substances.LT 4.9 – I can hypothesize and calculate the percent yield for a given chemical reaction. |
| ***Essential Question(s):*** Why does soda fizz?  |
| **Assessment**: * Complex stoichiometry exit ticket
* Collaborative Poster Rubric
* Percent Yield Exit Ticket
 |
| * ***Do Now***:
1. Draw and label what salt dissolving in water looks like
2. What is the molar mass of H to H2O
3. Given the following balanced equation: 8H + 2O2 🡪 4H2O
4. How many grams of Cl are produced from Salt given the following balanced equation: 4NaCl 🡪 4Na +2Cl2
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| **WHOLE GROUP/ DIRECT INTRUCTION** |
| * Complex Stoichiometry grams to grams
* Mole Ratios
* Percent Yield
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| **SMALL GROUP STATION** |  | **COLLABORATIVE STATION** |  | **COMPUTER ASSISTED STATION** |
| Practice Unit 3 moles to grams and grams to moles  |  | Around the world stations questions  |  | Notes on grams to grams complex stoichiometry  |