Thermodynamics

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| ***Teacher:*** | Ms. Athwal | ***Date:*** | March 24 – March 28 | ***Course:*** | Chemistry | ***Grade:*** | 11 |
| ***CA Standard(s):***  7c *Students know* energy is released when a material condenses or freezes and is absorbed when a material evaporates or melts.  7d *Students know* how to solve problems involving heat flow and temperature changes, using known values of specific heat and latent heat of phase change. | | | | | | | |
| ***Learning Objective (s):***  7.3 – I can draw a phase diagram to describe how pressure and temperature are related and discuss what a triple point is.  7.4 – I can analyze and draw a heating curve to describe how energy is lost or gained while a substance changes phases. I can identify on the graph the ΔHfusion and the ΔHvaporization.  7.5 – I can describe a reaction using the ideas of enthalpy and entropy. I can articulate how enthalpy and phase diagrams are related. | | | | | | | |
| * ***Essential Question(s):*** Why are flamin’ hot Cheetos so bad for me? ; How much energy is required to turn my Flamin’ Hot Cheetos into a liquid? | | | | | | | |
| **Assessment**:   * Literacy reading annotation * Exit ticket (3 questions on phase diagrams) * Quarter 3 Benchmarsk | | | | | | | |
| ***Do Now***:   1. Define endothermic. Give an example. 2. Define exothermic. Give an example 3. I cool a glass of water, is this an endothermic process or an exothermic process? | | | | | | | |

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| **WHOLE GROUP/ DIRECT INTRUCTION** |
| * Specific heat diagrams |

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| **SMALL GROUP STATION** |  | **COLLABORATIVE STATION** |  | **COMPUTER ASSISTED STATION** |
| Notes on **enthalpy** and **entropy**, and heating curves |  | Heating curve collaborative poster: With your table, complete the heating curve challenge problems.   * **Explain your reasoning next to each of your answers**     Reading high fructose corn syrup and food addiction articles |  | Benchmark 3 Reflection  Blogging Learning taret 7.1 and 7.2 |