Chemical Bonding

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| ***Teacher:*** | Ms. Athwal  | ***Date:*** | September 23-27 | ***Course:*** | Chemistry | ***Grade:*** | 11 |
| ***CA Standard(s):*** 2a *Students know* atoms combine to form molecules by sharing electrons to form covalent or metallic bonds or by exchanging electrons to form ionic bonds.2e *Students know* how to draw Lewis dot structures**.**2b *Students know* chemical bonds between atoms in molecules such as H2, CH4, NH3, H2CCH2, N2, Cl2, and many large biological molecules are covalent. |
| ***Learning Objective (s):*** LT 2.7 – For a given covalent compound, I can draw a Lewis Dot diagram for the compound and explain how it illustrates the covalent bond.LT 2.8 – I can analyze the similarities and differences between metallic, covalent, and ionic bonds. |
| ***Essential Question(s):*** What are conflict diamonds and what is the chemistry of the bonds in these diamonds? |
| **Assessment**: 1. Draw the Lewis structure for H2O
2. Draw the Lewis structure for NH3
3. Draw the Lewis structure for HCN
* Unit 2 Exam 9/26-9/27
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| * ***Do Now***:
* What is the difference between and ionic, covalent, and metallic bond
* What is the name of the force that holds crystal lattices together?
* Draw the Lewis Dot Structure for Ca
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| **WHOLE GROUP/ DIRECT INTRUCTION** |
| * Lewis Dot Structure for bonds
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| **SMALL GROUP STATION** |  | **COLLABORATIVE STATION** |  | **COMPUTER ASSISTED STATION** |
| Reviewing any concepts from Unit 2 that students are struggling with  |  | Unit 2 Stations Review  |  | Unit 2 Blogging  |